

Determination of Free Cyanide in a Copper Bath

Cyanide content determination in e.g. brass electroplating bath is very important to achieve optimum plating efficiency. Free cyanide is precipitated by silver nitrate, and the titration is monitored with a DP550 phototrode.

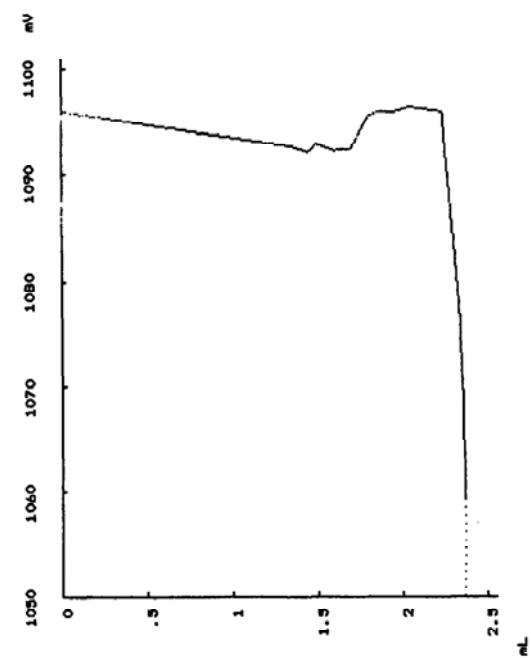
Sample	Copper bath, aliquots of 1 mL from diluted solution of 10 mL in 100 mL.	Preparation and Procedures
Substance	Free cyanide, CN^- , M = 26.02, $z = 1$	<p>Sample preparation</p> <ol style="list-style-type: none"> 10 mL bath are diluted to 100 mL with deion. water. 1 mL aliquot is poured in the titration beaker. Add 5 mL 10% potassium iodide solution (KI). 40 mL deionized water is added to the sample. <p>A soluble complex $[\text{Ag}(\text{CN})_2]^-$ is first formed:</p> $\text{Ag}^+ + 2 \text{CN}^- = [\text{Ag}(\text{CN})_2]^-$ <p>As long as free cyanide is still present, the solution remains clear, but the first excess of silver causes formation of a white solid that mark the endpoint:</p> $\text{Ag}^+ + \text{Ag}(\text{CN})_2^- = \text{Ag}[\text{Ag}(\text{CN})_2]$ <p>For an accurate end-point determination, KI is used as an indicator. During titration, any silver iodide which would tend to form will be kept in solution by the excess of cyanide ion until the equivalence point is reached:</p> $\text{AgI} + 2 \text{CN}^- = [\text{Ag}(\text{CN})_2]^- + \text{I}^-$ <p>After the EQP, the formation of yellow silver iodide will lead to a turbidity increase which indicates the endpoint.</p>
Chemicals	5 mL potassium iodide, KI, 10% 40 mL deion. water	
Titrant	Silver nitrate, AgNO_3 $c(1/2 \text{AgNO}_3) = 0.2 \text{ mol/L}$	
Standard	Sodium chloride (see appl. M525)	
Instruments	DL50 Graphix, DL53/DL55/DL58, DL70ES/DL77 AT261, Printer	
Accessories	Titration beakers ME-101974	
Indication	DP550 Phototrode with DIN-Lemo Adapter ME-89600	<p>Remarks</p> <p>The method was developed on the DL25 titrator and has been adapted for the DL5x-/DL7x-titrators.</p>
Chemistry	$\text{Ag}^+ + 2 \text{CN}^- = \text{Ag}(\text{CN})_2^-$ $\text{Ag}^+ + \text{Ag}(\text{CN})_2^- = \text{Ag}[\text{Ag}(\text{CN})_2]$ $\text{AgI} + 2 \text{CN}^- = [\text{Ag}(\text{CN})_2]^- + \text{I}^-$	<p>Chemicals</p> <p>10% Potassium iodide, KI: 50 g KI and 50 g NaOH are dissolved in deionized water, and diluted to 500 mL in a volumetric flask.</p> <p>CAUTION: Cyanide is toxic. Wear safety goggles and gloves, and work in a fume hood. NEVER add strong acid to the solution since cyanidic acid is formed and liberated as a gas from the solution.</p>
Calculation	$R = Q \cdot C/m$ $C = M/z$ $R2 = R1 \cdot 10$ $R3 = Q \cdot C3$ $C3 = M \cdot 1000/z$	<p>Literature</p> <ol style="list-style-type: none"> 1. Application note, DL25 Application Brochure "Petroleum and electroplating", ME-51724627. 2. Vogel's textbook of quantitative inorganic analysis, 4th edition, Longman Group Limited, 1978. 3. D.A. Skoog, D.M. West, "Fundamentals of Analytical Chemistry", Holt, Rinehart, and Winston, 1969. 4. Application no. M525 in Application brochure 18, "Standardization of Titrants", 2000.
Waste disposal	Cyanide waste. CAUTION: cyanide is toxic.	
Author	R. Böhlen, MT-CH	

Results

	CN-	n	Comments
Mean value	26.84 g/L	6	DL25 application note in
Standard deviation s	0.346 g/L		DL25 Application brochure "Petroleum products and electroplating" ME-51 724 627
Rel. standard deviation srel	1.29 %		

Table of measured values

Titration curve



Method

Method	25002	Free CN- in Cu bath
Version	17-01-2001	10:50
Title		
Method ID	25002	
Title		Free CN- in Cu bath
Date/time		17-01-2001 10:50
Sample		
Sample ID		Cu bath
Entry type		Fixed volume
Volume [mL]		1.0
Molar mass M		26.01
Equivalent number z		1
Titration stand		Stand 1
Temperature sensor		Manual
Stir		
Speed [%]		50
Time [s]		10
EQP titration		
Titrant/Sensor		
Titrant		1/2AgNO ₃
Concentration [mol/L]		0.2
Sensor		DP550
Unit of meas.		mV
Predispensing		to volume
Volume [mL]		1.5
Wait time [s]		20
Titrant addition		Dynamic
dE(set) [mV]		4.0
dV(min) [mL]		0.02
dV(max) [mL]		0.2
Measure mode		Equilibrium controlled
dE [mV]		0.5
dt [s]		1.0
t(min) [s]		5.0
t(max) [s]		30.0
Recognition		
Threshold		200
Steepest jump only		No
Range		No
Tendency		Negative
Termination		
at maximum volume [mL]		20.0
at potential		No
at slope		No
after number EQPs		Yes
n =		1
comb. termination criteria		No
Evaluation		
Procedure		Standard
Potential 1		No
Potential 2		No
Stop for reevaluation		Yes
Condition		neq=0
Calculation		
Formula		R=Q*C/m
Constant		C=M/z
Decimal places		2
Result unit		g/L
Result name		Sample content
Statistics		Yes
Calculation		
Formula		R2=R1*10
Constant		
Decimal places		2
Result unit		g/L
Result name		Bath content
Statistics		Yes
Calculation		
Formula		R3=Q*C3
Constant		C3=M*1000/z
Decimal places		2
Result unit		mg
Result name		CN in sample
Statistics		No
Report		
Output		Printer
Results		No
All results		Yes
Raw results		No
Table of measured values		Yes
Sample data		No
E - V curve		Yes
dE/dV - V curve		Yes
d2E/dV ² - V curve		No
log dE/dV - V curve		No
E - t curve		No
V - t curve		No
dV/dt - t curve		No