

Determination of Free Cyanide in a Copper Bath

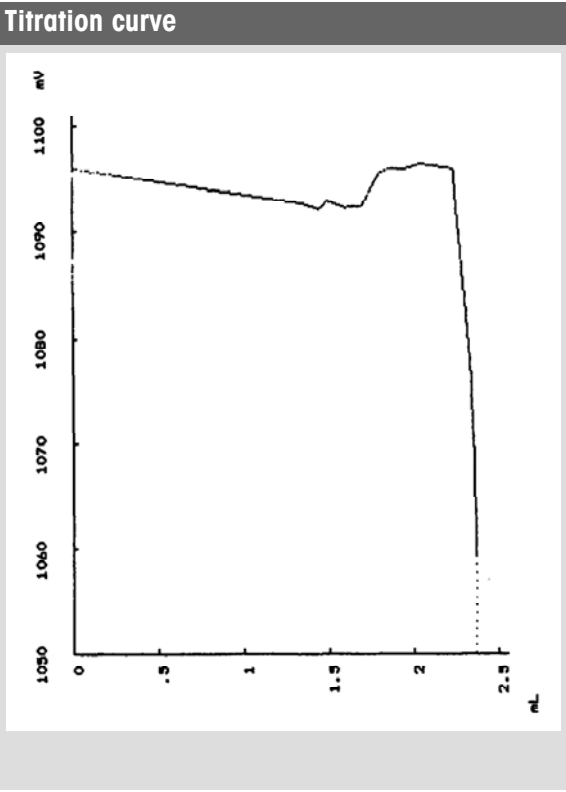
Cyanide content determination in e.g. brass electroplating bath is very important to achieve optimum plating efficiency. Free cyanide is precipitated by silver nitrate, and the titration is monitored with a DP550 phototrode.

Sample	Copper bath, aliquots of 1 mL from diluted solution of 10 mL in 100 mL.	Preparation and Procedures Sample preparation 1. 10 mL bath are diluted to 100 mL with deion. water. 2. 1 mL aliquot is poured in the titration beaker. 3. Add 5 mL 10% potassium iodide solution (KI). 4. 40 mL deionized water is added to the sample. A soluble complex $[\text{Ag}(\text{CN})_2]^-$ is first formed: $\text{Ag}^+ + 2 \text{CN}^- = [\text{Ag}(\text{CN})_2]^-$ As long as free cyanide is still present, the solution remains clear, but the first excess of silver causes formation of a white solid that mark the endpoint: $\text{Ag}^+ + \text{Ag}(\text{CN})_2^- = \text{Ag}[\text{Ag}(\text{CN})_2]$ For an accurate end-point determination, KI is used as an indicator. During titration, any silver iodide which would tend to form will be kept in solution by the excess of cyanide ion until the equivalence point is reached: $\text{AgI} + 2 \text{CN}^- = [\text{Ag}(\text{CN})_2]^- + \text{I}^-$ After the EQP, the formation of yellow silver iodide will lead to a turbidity increase which indicates the endpoint.
Substance	Free cyanide, CN^- , $M = 26.02$, $z = 1$	
Chemicals	5 mL potassium iodide, KI, 10% 40 mL deion. water	
Titrant	Silver nitrate, AgNO_3 $c(1/2 \text{ AgNO}_3) = 0.2 \text{ mol/L}$	
Standard	Sodium chloride (see appl. M525)	
Instruments	DL50 Graphix, DL53/DL55/DL58, DL70ES/DL77 AT261, Printer	
Accessories	Titration beakers ME-101974	Remarks The method was developed on the DL25 titrator and has been adapted for the DL5x-/DL7x-titrators. Chemicals 10% Potassium iodide, KI: 50 g KI and 50 g NaOH are dissolved in deionized water, and diluted to 500 mL in a volumetric flask. CAUTION: Cyanide is toxic. Wear safety goggles and gloves, and work in a fume hood. NEVER add strong acid to the solution since cyanidric acid is formed and liberated as a gas from the solution. Literature 1. Application note, DL25 Application Brochure "Petroleum and electroplating", ME-51724627. 2. Vogel's textbook of quantitative inorganic analysis, 4th edition, Longman Group Limited, 1978. 3. D.A. Skoog, D.M. West, "Fundamentals of Analytical Chemistry", Holt, Rinehart, and Winston, 1969. 4. Application no. M525 in Application brochure18, "Standardization of Titrants", 2000.
Indication	DP550 Phototrode with DIN-Lemo Adapter ME-89600	
Chemistry	$\text{Ag}^+ + 2 \text{CN}^- = \text{Ag}(\text{CN})_2^-$ $\text{Ag}^+ + \text{Ag}(\text{CN})_2^- = \text{Ag}[\text{Ag}(\text{CN})_2]$ $\text{AgI} + 2 \text{CN}^- = [\text{Ag}(\text{CN})_2]^- + \text{I}^-$	
Calculation	$R=Q \cdot C/m \quad ; \quad \text{CN in sample (g/L)}$ $C=M/z$ $R2=R1 \cdot 10 \quad ; \quad \text{CN Bath content in g/L}$ $R3=Q \cdot C3 \quad ; \quad \text{CN in sample (mg)}$ $C3=M \cdot 1000/z$	
Waste disposal	Cyanide waste. CAUTION: cyanide is toxic.	
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Results

	CN-	n	Comments
Mean value	26.84 g/L	6	DL25 application note in
Standard deviation s	0.346 g/L		DL25 Application brochure "Petroleum products and electroplating" ME-51 724 627
Rel. standard deviation srel	1.29 %		

Table of measured values



Method

Method	25002	Free CN- in Cu bath
Version	17-01-2001	10:50
Title		
Method ID	25002	
Title	Free CN- in Cu bath	
Date/time	17-01-2001	10:50
Sample		
Sample ID	Cu bath	
Entry type	Fixed volume	
Volume [mL]	1.0	
Molar mass M	26.01	
Equivalent number z	1	
Titration stand	Stand 1	
Temperature sensor	Manual	
Stir		
Speed [%]	50	
Time [s]	10	
EQP titration		
 Titrant/Sensor		
Titrant	1/2AgNO3	
Concentration [mol/L]	0.2	
Sensor	DP550	
Unit of meas.	mV	
 Predispensing		
to volume		
Volume [mL]	1.5	
Wait time [s]	20	
 Titrant addition		
Dynamic		
dE(set) [mV]	4.0	
dV(min) [mL]	0.02	
dV(max) [mL]	0.2	
 Measure mode		
Equilibrium controlled		
dE [mV]	0.5	
dt [s]	1.0	
t(min) [s]	5.0	
t(max) [s]	30.0	
 Recognition		
Threshold	200	
Steepest jump only	No	
Range	No	
Tendency	Negative	
 Termination		
at maximum volume [mL]	20.0	
at potential	No	
at slope	No	
after number EQPs	Yes	
n =	1	
comb. termination criteria	No	
 Evaluation		
Procedure	Standard	
Potential 1	No	
Potential 2	No	
Stop for reevaluation	Yes	
Condition	neq=0	
Calculation		
Formula	R=Q*C/m	
Constant	C=M/z	
Decimal places	2	
Result unit	g/L	
Result name	Sample content	
Statistics	Yes	
Calculation		
Formula	R2=R1*10	
Constant		
Decimal places	2	
Result unit	g/L	
Result name	Bath content	
Statistics	Yes	
Calculation		
Formula	R3=Q*C3	
Constant	C3=M*1000/z	
Decimal places	2	
Result unit	mg	
Result name	CN in sample	
Statistics	No	
Report		
Output	Printer	
Results	No	
All results	Yes	
Raw results	No	
Table of measured values	Yes	
Sample data	No	
E - V curve	Yes	
dE/dV - V curve	Yes	
d2E/dV2 - V curve	No	
log dE/dV - V curve	No	
E - t curve	No	
V - t curve	No	
dV/dt - t curve	No	